

Lignin-based binder for high performance carbon electrode

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Concept of lignin-binders

Polyvinylidene fluoride (PVDF) is the most conventional binder used in commercial Li-ion electrodes. However, PVDF binder is expensive, not easily to recycle and it involves the usage of volatile organic compounds for its processing. In addition, the reaction between PVDF with Li is an exothermal, which can cause self-heating and thermal runaway.

Therefore, Lignin-based binder is studied as an alternative binders that are more environmentally friendly. Lignin has emerged as an interesting raw material for the carbon materials in an electrode due to its low price and resistance to heat. Besides that, it can modify the electrode due to the formation of electroactive quinone functionalities which can oxidize and reduce during cycling.

Results

1. Stereomicroscope images



Enzymatic hydrolysed wheat straw

2. SEM images



Lignin reacted with the solvent



Acetylene carbon black as active material

mats



Untreated carbon from 600°C pyrolysis as active material

structure.



Schematic of conventional positive electrode for lithium ion battery

Structure of three phenyl propane monomers in lignin.



Experiment Methods

- Stereomicroscope a)
- b) Scanning electron microscope (SEM)
- X-Ray diffraction (XRD) measurement C)





3. XRD Analysis

destruction of fibers

Source: Dorner-Reisel. A.: Examination report XRD for wheat straw lignins of PFI e.V. Hochschule Schmalkalden. 31.08.2017





		ISIT)
45000 21	COMPARISON BETWEEN BIOMASS 0, 2.0 AND 2.1-PFI	VTEN
40000 -	21	≤
35000 -	Biomass 0	
20000	1 1 1 1 Biomass 2.0	
50000	$\mathbb{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}_{1}] = \mathbb{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal{E}[\mathcal$	

Crl value increased after undergo thermal pressure hydrolysis. Reason: Removal of non-cellulosic amorphous polysaccharides and less ordered surface cellulose chains in the matrix.

Crl value decreases after enzymatic hydrolysis. Reason: Breaking down of the cellulose molecular chain into sugar, which gradually decrease the degree of polymerization in cellulose.

Shifting (~ 0.50° to 1.00°) towards higher 20 Reason: Crystals dilation.

4. Raman Analysis



Comparison between the electrodes, that use acetylene carbon black as active material. As temperature gets higher, the intensity gets lower. Reason :

- Dispersing of graphene by NMP and Acetone.
- The bundles formed were more loosely bundled at high

Source: Dorner-Reisel. A.: Examination report XRD for wheat straw lignins of PFI e.V. Hochschule Schmalkalden. 31.08.2017

temperature.

Reduction in the degree of intermolecular association occuring within the lignin.

Raman spectroscopy d) Resistivity of electrodes e)





Arrangement of electrodes on the square steel plates

References

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Resistivity of mixture C and D (Ωm)				Resistivity of mixture E (Ω m)			
350	Mixture C	Mixture D	-Mixture E	8000			
300 -	$\mathbf{\lambda}$			- 7000			
250 -				- 6000			
200 -				- 5000			
				- 4000			
150 -				- 3000			
100 -		\setminus /		- 2000			
50 -		$ \rightarrow $		- 1000			
0	50			0			
	50	90	110				
Temperature (°C)							

Mixture E (Untreated carbon from 600°C pyrolysis as active material) has highest resistance. Reason :

- Low crystallinity of the 600°C pyrolysis
- High percentage of defects on the graphene surfaces
- Low contact between the carbon particles

Resistance of mixture C, which contain lignin and solvent, decreases as temperature increases. Reason :

Reducing of the glass transition temperature, T_g . Lignin are converted into a rubber-like state and it is more elastic

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